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**NATIONAL STANDARD OF THE
PEOPLE'S REPUBLIC OF CHINA**

GB/T 9737-2008

Replacing GB/T 9737-1988

**Chemical reagent -- General rules for the determination of
readily carbonizable substances**

化学试剂 易炭化物质测定通则

(English Translation)

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Foreword

This standard and ISO 6353-1. 1982 "Chemical analysis of reagent --- Part 1. General test method" non-equivalent.

This standard replaces GB/T 9737-1988 "Determination of chemical reagents easily carbonized material General", compared with the GB/T 9737-1988 Main Changes are as follows.

--- Increasing the cobalt chloride solution, a solution of copper sulfate, ferric chloride solution shelf (Excerpts 4.3,4.4,4.5);

--- Modify the mass fraction of cobalt chloride solution concentration calibration and calibration concentration copper sulfate solution of acetic acid solution (1988 Version of Appendix A; this edition of Appendix A);

--- Canceled Appendix B (1988 edition).

Appendix A of this standard is a normative appendix.

The standard proposed by China Petroleum and Chemical Industry Association.

This standard by the National Standardization Technical Committee Chemistry Chemicals Branch (SAC/TC63/SC3) centralized.

This standard was drafted. Sinopharm Chemical Reagent Co., Ltd.

1 Scope

This standard specifies 20 different shades of color standards, and be easy to use these standard colors carbonized substance determined General.

This standard applies to the determination of organic chemicals easily carbonized

substances.

2 Normative references

The following documents contain provisions which, through reference in this standard and become the standard terms. For dated references, subsequent Amendments (not including errata content) or revisions do not apply to this standard, however, encourage the parties to the agreement are based on research Whether the latest versions of these documents. For undated reference documents, the latest versions apply to this standard.

Preparation of GB/T 601 chemical reagent standard titration solution GB/T 603 chemical reagent test method Preparations of articles (GB/T 603-2002, ISO 6353-1 with. 1982, NEQ) GB/T 6682 analytical laboratory use specifications and test methods (GB/T 6682-2008, ISO 3696. 1987, MOD)

3 PRINCIPLE OF THE METHOD

Concentrated sulfuric acid composition of the water can be captured hydrogen and oxygen from some organic compounds, the carbonization of these organic compounds, so that the concentrated sulfuric acid or Mixture of concentrated sulfuric acid and the sample color rendering, can be used to visually easily carbonized material colorimetry.

4.1 General provisions

Unless otherwise specified in this standard, the standard titration solution used in formulations and products, according to GB/T 601, the provisions of the preparation of GB/T 603's;

Test water should comply with GB/T 6682 water in three specifications, are used in mass fraction in% of the solution.

4.2 Hydrochloric acid solution (39 + 1)

Measure 3900mL of water, was added 100mL of hydrochloric acid, and mix.

Weigh 60g cobalt chloride hexahydrate, was dissolved in 900mL of hydrochloric acid solution (39 + 1) and diluted with hydrochloric acid (39 + 1) to 1000mL. According to the method specified in Appendix A calibration, and by calculating the amount of added hydrochloric acid solution (39 + 1), the final mass of the solution Concentration of 59.5mg/mL (in $\text{CoCl}_2 \cdot 6\text{H}_2\text{O}$ dollars). The solution was kept in the dark sealed package, valid for six months.

Weigh 63g copper sulfate pentahydrate, dissolved in 900mL of hydrochloric acid solution (39 + 1) and diluted with hydrochloric acid (39 + 1) to 1000mL. According to the method specified in Appendix A calibration, and by calculating the amount of added hydrochloric

acid solution (39 + 1), the final concentration of the solution it was 62.4mg/mL (in $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$ dollars). The solution was kept in the dark sealed package, valid for six months.

Weigh 46g of ferric chloride ($\text{FeCl}_3 \cdot 6\text{H}_2\text{O}$), was dissolved in 900mL of hydrochloric acid solution (39 + 1) and (1 + 39) was diluted with hydrochloric acid to 1000mL. According to the method specified in Appendix A calibration, and by calculating the amount of added hydrochloric acid solution (39 + 1), the final concentration of the solution Degree of 45.0mg/mL ($\text{FeCl}_3 \cdot 6\text{H}_2\text{O}$ to count). The solution was kept in the dark sealed package, valid for six months.